Chemistry – Atomic Structure

Q. 1 Match the correct pairs : (Use the pencil and draw lines)

Element [protons]		Electronic Configuration
1. Magnesium[p=12]	a)	2, 8, 3
2. Sulphur[p=16]	b)	2, 8
3. Neon[p=10]	c)	2, 8, 2
4. Calcium[p=20]	d)	2, 8, 6
5. Aluminium[p=13]	e)	2, 8, 8, 2

Q. 2 Choose the word options and drag them into the correct places:

$2n^2$ | Mass No. | Isotopes | Valency | Atomic No.

1. Atoms of the same element having the atomic number but different mass number. _____
2. This is to tell us the total number of protons and neutrons present in a neutral atom. _____
3. It is formula to determine the maximum number of electrons in each shell of an atom. _____
4. This is to tell us the number of electrons or protons that are present in a neutral atom. _____
5. This tells us the number of electrons donated or accepted by an atom to achieve a stable electronic configuration. _____

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Q. 3 Drag the electron/s from the first figure and drop it into the second figure to show the electron transfer and then answer the questions given below:

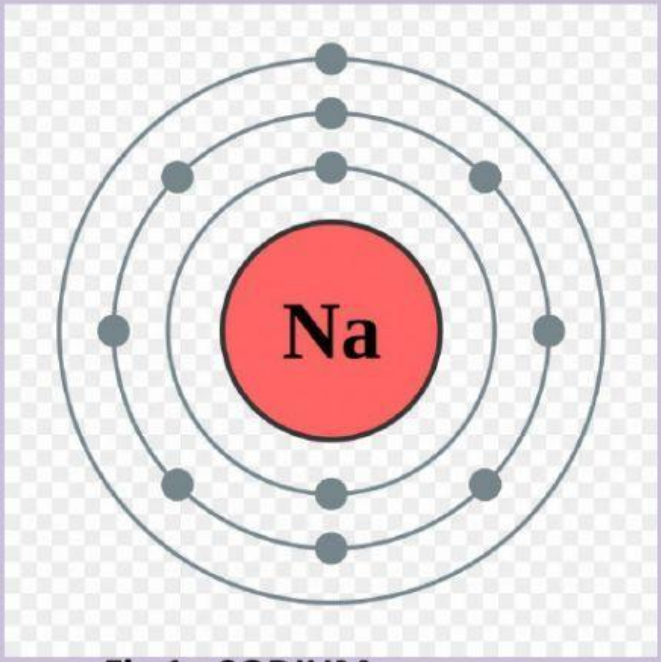

Fig 1 : SODIUM

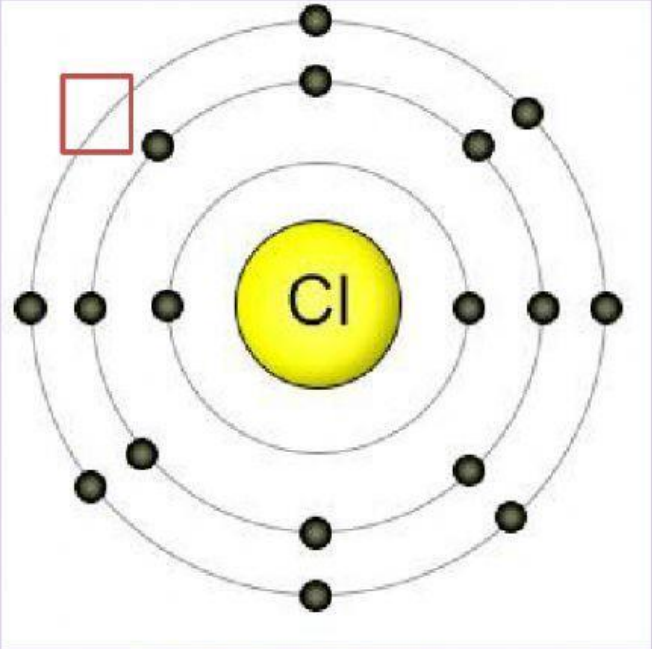

Fig 2 : CHLORINE

1. What kind of bond is achieved by this method? .....
2. Which of these two is a metal? .....
3. Which is the nearest inert gas to Chlorine? .....
4. Write the electronic configuration of Sodium. .....
5. What is the valency of Chlorine? .....

Q. 4 Match the element symbol with its valency : (Draw lines)

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