

3. Substance: Carbon dioxide

Phase: Solid

Phase: Gas

Volume

Mass

Density

100 cm³

156 g

_____ g/cm³
answer
2 decimals

Volume

Mass

100 cm³

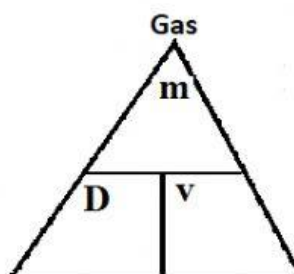
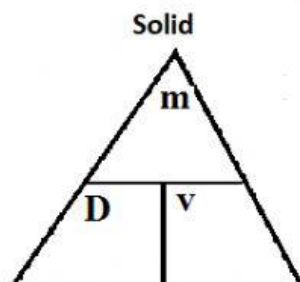
0.198 g
4 decimals

Density

g/cm³

answer
5 decimals

Name of phase change from solid to gas _____



Which is more dense, the solid or gas? **solid gas**

How does the arrangement of the particles in a solid and a gas explain your answer?

The particles in a **solid gas** are packed more closely together this **increases decreases** density.

4. Substance: Mercury

Phase: Solid

Phase: Liquid

Volume

Mass

Density

10 cm³

136.5 g

_____ g/cm³
answer
2 decimals

Volume

Mass

10 cm³

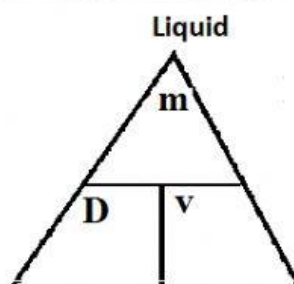
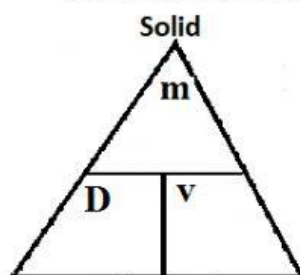
135.9 g

Density

g/cm³

answer
2 decimals

Name of phase change from liquid to solid _____



Which is more dense, the solid or liquid? **solid liquid**

How does the arrangement of the particles in a solid and a liquid explain your answer?

The particles in a **solid liquid** are packed more closely together this **increases decreases** density.

5. Substance: Pentane

Phase: Liquid

Phase: Gas

Volume

Mass

Density

3800 cm³

2370 g

_____ g/cm³
answer
2 decimals

Volume

Mass

3800 cm³

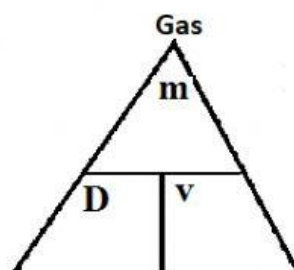
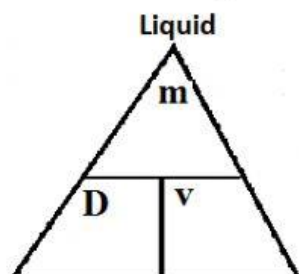
10.7 g

Density

g/cm³

answer
4 decimals

Name of phase change from gas to liquid _____



Which is more dense, the liquid or gas? **liquid gas**

How does the arrangement of the particles in a liquid and a gas explain your answer?

The particles in a **liquid gas** are packed more closely together this **increases decreases** density.